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1. 88% The actual yield of a reaction is 22 g and the theoretical

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1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g.

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Vocabulary. STUDY. PLAY. Limiting

reactant. The reactant that limits

the amount of the other reactant

that can combine and the amount

of product that can form in a

chemical reaction. Excess reactant.

The substance that is not used up

completely in a reaction. Holt

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PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The

actual yield of a reaction is 22 g and the theoretical yield is 25 g.

Calculate the percentage yield. 2.

6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the

following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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vi Chemistry Section 5 1 Review Answer. Holt Modern Chemistry Chapter 9 Review Stoichiometry. Chemical Equations and Reactions SECTION 82 SHORT ANSWER

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